FIITJEE

JEE(Main)- 2017

CHEMISTRY

PHASE TEST – IV

Time Allotted: 1 Hr. Maximum Marks:144

- Please read the instructions carefully. You are allotted 5 minutes specifically for this purpose.
- You are not allowed to leave the Examination Hall before the end of the test.

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INSTRUCTIONS

A. General Instructions

1. This booklet is your question paper. Answers have to be marked on the provided OMR sheets.
2. This question paper contains 30 questions. All questions are compulsory.
3. Rough spaces are provided for rough work inside the question paper. No additional sheets will be provided for rough work.
4. Blank Papers, clip boards, log tables, slide rule, calculator, cellular phones, pagers and electronic devices, in any form, are not allowed.
5. Write in your Name and the Enrolment No. in the space provided at the bottom of this page.

B. Filling of OMR Sheet

6. On the OMR sheet, write in ink your Name, Enrolment No., and Name of the centre and put your signature in the appropriate boxes.
7. Every question has four choices for its answers (A), (B), (C), (D)

C. Marking Scheme

8. Questions **1 to 03** have only one correct answer and carries **+8 marks** each for correct answer and **– 2 mark** for each wrong answer.
9. Questions **4 to 09** have only one correct answer and carries **+4 marks** each for correct answer and **– 1 mark** for each wrong answer.
10. Questions **10 to 12** have only one correct answer and carries **+8 marks** each for correct answer and **– 2 mark** for each wrong answer.
11. Questions **13 to 30** have only one correct answer and carries **+4 marks** each for correct answer and **– 1 mark** for each wrong answer.

Name of the Candidate ☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐☐

Enrolment No. ☐☐☐☐☐☐☐☐☐☐☐☐☐☐

RANKERS STUDY MATERIAL

Useful Data

Gas Constant	**R**	=	8.314 J K^{-1} mol^{-1}
		=	0.0821 Lit atm K^{-1} mol^{-1}
		=	1.987 ≈ 2 Cal K^{-1} mol^{-1}
Avogadro's Number	N_a	=	6.023×10^{23}
Planck's constant	**h**	=	6.625×10^{-34} J·s
		=	6.625×10^{-27} erg·s
1 Faraday		=	96500 coulomb
1 calorie		=	4.2 joule
1 amu		**=**	1.66×10^{-27} kg
1 eV		**=**	1.6×10^{-19} J

Atomic No: H=1, He = 2, Li=3, Be=4, B=5, C=6, N=7, O=8, N=9, Na=11, Mg=12, Si=14, Al=13, P=15, S=16, Cl=17, Ar=18, K =19, Ca=20, Cr=24, Mn=25, Fe=26, Co=27, Ni=28, Cu = 29, Zn=30, As=33, Br=35, Ag=47, Sn=50, I=53, Xe=54, Ba=56, Pb=82, U=92.

Atomic masses: H=1, He=4, Li=7, Be=9, B=11, C=12, N=14, O=16, F=19, Na=23, Mg=24, Al = 27, Si=28, P=31, S=32, Cl=35.5, K=39, Ca=40, Cr=52, Mn=55, Fe=56, Co=59, Ni=58.7, Cu=63.5, Zn=65.4, As=75, Br=80, Ag=108, Sn=118.7, I=127, Xe=131, Ba=137, Pb=207, U=238.